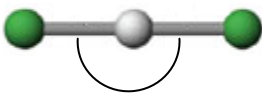
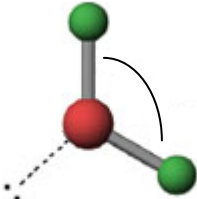
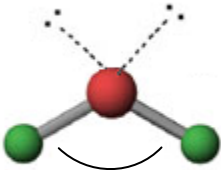
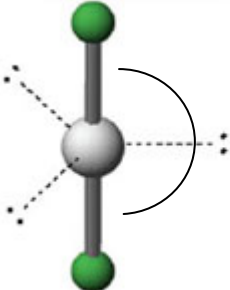
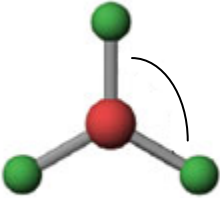
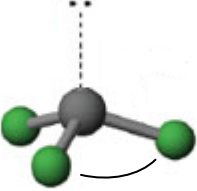
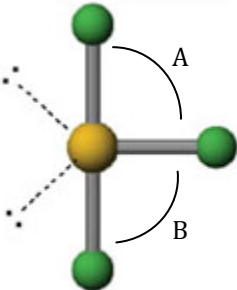
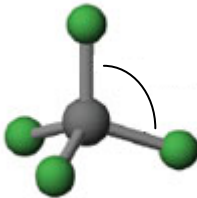


## VSEPR GEOMETRIES

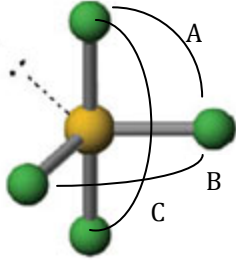
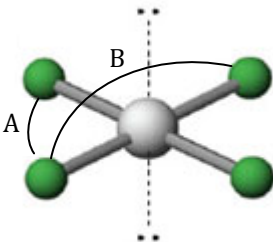
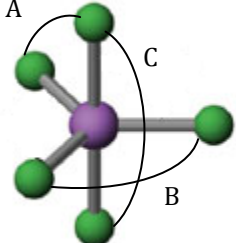
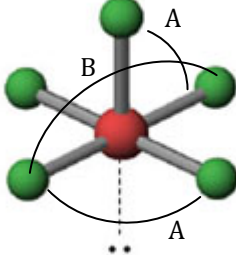
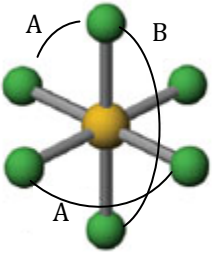
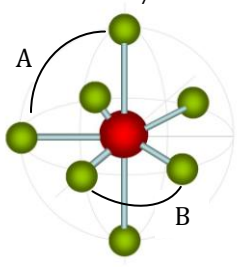
VSEPR Type	Bond Angle(s)	Polar   Nonpolar	VSEPR Type	Bond Angle(s)	Polar   Nonpolar
$AX_2$  Linear	$180^\circ$	$N_2O$   $CO_2$	$AX_2E$  Bend/Angular	$120^\circ$	$O_3$   NONE
$AX_2E_2$  Bent/Angular	$104.5^\circ$	$H_2O$   NONE	$AX_2E_3$  Linear	$180^\circ$	$BrF_3^-$   $XeF_2$
$AX_3$  Triangular Planar	$120^\circ$	$COCl_2$   $BF_3$	$AX_3E$  Triangular Pyramidal	$107.3^\circ$	$NH_3$   NONE
$AX_3E_2$  T-Shaped	$A/B: 90^\circ$ $A + B: 180^\circ$	$CIF_3$   NONE	$AX_4$  Tetrahedral	$109.5^\circ$	$CH_3O$   $CH_4$

A = Central Atom

X = Number of Terminal Atoms

E = Number of lone pairs on the central atom (A)

## VSEPR GEOMETRIES

VSEPR Type	Bond Angle(s)	Polar   Nonpolar	VSEPR Type	Bond Angle(s)	Polar   Nonpolar
$AX_4E$  Seesaw	A: $90^\circ$ B: $120^\circ$ C: $180^\circ$	SF <sub>4</sub>   NONE	$AX_4E_2$  Square Planar	A: $90^\circ$ B: $180^\circ$	XeF <sub>3</sub> Br   RnF <sub>4</sub>
$AX_5$  Triangular Bi-pyramidal	A: $90^\circ$ B: $120^\circ$ C: $180^\circ$	PCl <sub>4</sub> F   PCl <sub>5</sub>	$AX_5E$  Square Pyramidal	A: $90^\circ$ B: $180^\circ$	IF <sub>5</sub>   NONE
$AX_6$  Octahedral	A: $90^\circ$ B: $180^\circ$	IOF <sub>5</sub>   SF <sub>6</sub>	$AX_7$  Pentagonal Bi-pyramidal	A: $90^\circ$ B: $72^\circ$	NONE   IF <sub>7</sub>

A = Central Atom

X = Number of Terminal Atoms

E = Number of lone pairs on the central atom (A)